

The oligonucleotide-modified nanoparticles are stable at elevated temperatures (80°C) and high salt concentrations (1M NaCl) for days and have not been observed to undergo particle growth. Stability in high salt concentrations is important, since such conditions are required for the hybridization reactions that form the basis of the methods of detection and nanofabrication of the invention.

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Example 2: Formation Of Nanoparticle Aggregates

A. Preparation Of Linking Oligonucleotide

Two (nonthiolated) oligonucleotides were synthesized as described in part B of Example 1. They had the following sequences:

10 3' ATATGCGCGA TCTCAGCAAA [SEQ ID NO:1]; and

3' GATCGCGCAT ATCAACGGTA [SEQ ID NO:2].

15 Mixing of these two oligonucleotides in a 1 M NaCl, 10 mM phosphate buffered (pH 7.0) solution, resulted in hybridization to form a duplex having a 12-base-pair overlap and two 8-base-pair sticky ends. Each of the sticky ends had a sequence which was complementary to that of one of the oligonucleotides attached to the Au colloids prepared in part C of Example 1.

B. Formation Of Nanoparticle Aggregates

20 The linking oligonucleotides prepared in part A of this example (0.17 μ M final concentration after dilution with NaCl) were added to the nanoparticle-oligonucleotide conjugates prepared in part C of Example 1 (5.1 nM final concentration after dilution with NaCl) at room temperature. The solution was then diluted with aqueous NaCl (to a final concentration of 1 M) and buffered at pH 7 with 10 mM phosphate, conditions which are suitable for hybridization of the oligonucleotides. An immediate color change from red to purple was observed, and a precipitation reaction ensued. See Figure 6. Over the course of 25 several hours, the solution became clear and a pinkish-gray precipitate settled to the bottom of the reaction vessel. See Figure 6.

100-1000-1000-1000-1000

To verify that this process involved both the oligonucleotides and colloids, the precipitate was collected and resuspended (by shaking) in 1 M aqueous NaCl buffered at pH 7. Any of the oligonucleotides not hybridized to the nanoparticles are removed in this manner. Then, a temperature/time dissociation experiment was performed by monitoring the characteristic absorbance for the hybridized oligodeoxyribonucleotides (260 nm) and for the aggregated colloids which is reflective of the gold interparticle distance (700 nm). See Figure 7.

Changes in absorbance at 260 and 700 nm were recorded on a Perkin-Elmer Lambda 2 UV-vis Spectrophotometer using a Peltier PTP-1 Temperature Controlled Cell Holder while cycling the temperature at a rate of 1°C/minute between 0°C and 80°C. DNA solutions were approximately 1 absorbance unit(s) (OD), buffered at pH 7 using 10 mM phosphate buffer and at 1M NaCl concentration.

The results are shown in Figure 8A. As the temperature was cycled between 0°C and 80°C (which is 38°C above the dissociation temperature (T_m) for the duplex ($T_m=42$ °C)), there was an excellent correlation between the optical signatures for both the colloids and oligonucleotides. The UV-vis spectrum for naked Au colloids was much less temperature dependent, Figure 8B.

There was a substantial visible optical change when the polymeric oligonucleotide-colloid precipitate was heated above its melting point. The clear solution turned dark red as the polymeric biomaterial de-hybridized to generate the unlinked colloids which are soluble in the aqueous solution. The process was reversible, as evidenced by the temperature traces in Figure 8A.

In a control experiment, a 14-T:14-A duplex was shown to be ineffective at inducing reversible Au colloid particle aggregation. In another control experiment, a linking oligonucleotide duplex with four base pair mismatches in the sticky ends was found not to induce reversible particle aggregation of oligonucleotide-modified nanoparticles (prepared as described in part C of Example 1 and reacted as described above). In a third control experiment, non-thiolated oligonucleotides having sequences complementary to the sticky

ends of the linking oligonucleotide and reacted with nanoparticles did not produce reversible aggregation when the nanoparticles were combined with the linking oligonucleotide.

Further evidence of the polymerization/assembly process came from Transmission Electron Microscopy (TEM) studies of the precipitate. TEM was performed on a Hitachi 8100 Transmission Electron Microscope. A typical sample was prepared by dropping 100 μ L of colloid solution onto a holey carbon grid. The grid, then, was dried under vacuum and imaged. TEM images of Au colloids linked by hybridized oligonucleotides showed large assembled networks of the Au colloids, Figure 9A. Naked Au colloids do not aggregate under comparable conditions but rather disperse or undergo particle growth reactions. Hayat, *Colloidal Gold: Principles, Methods, and Applications* (Academic Press, San Diego, 1991). Note that there is no evidence of colloid particle growth in the experiments performed to date; the hybridized colloids seem to be remarkably regular in size with an average diameter of 13 nm.

With TEM, a superposition of layers is obtained, making it difficult to assess the degree of order for three-dimensional aggregates. However, smaller scale images of single layer, two-dimensional aggregates provided more evidence for the self-assembly process, Figure 9B. Close-packed assemblies of the aggregates with uniform particle separations of approximately 60 \AA can be seen. This distance is somewhat shorter than the estimated 95 \AA spacing expected for colloids connected by rigid oligonucleotide hybrids with the sequences that were used. However, because of the nicks in the duplex obtained after hybridization of the oligonucleotides on the nanoparticles to the linking oligonucleotides, these were not rigid hybrids and were quite flexible. It should be noted that this is a variable that can be controlled by reducing the system from four overlapping strands to three (thereby reducing the number of nicks) or by using triplexes instead of duplexes.